

Balancing and Reaction Type PRACTICE

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Balancing chemical equations involves counting the atoms of each element and making sure that they are equal on both sides. If they are, the equation is said to be "balanced." If not, then the equation must be balanced **by adding coefficients only**. NEVER change the subscript number of a formula when balancing, this changes the compound.

There are four main types of reactions that have been discussed in class. Using the example provided for a synthesis reaction, write a brief description of what happens in each of the other types of reactions.

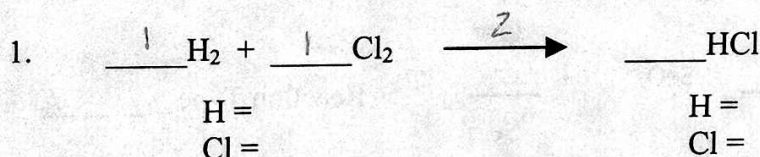
Synthesis: In a synthesis reaction two simple substances are coming together to form a more complex compound

Decomposition: _____

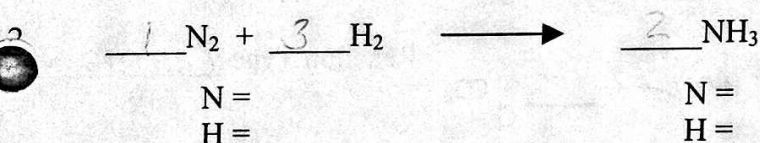
Single replacement: _____

Double replacement: _____

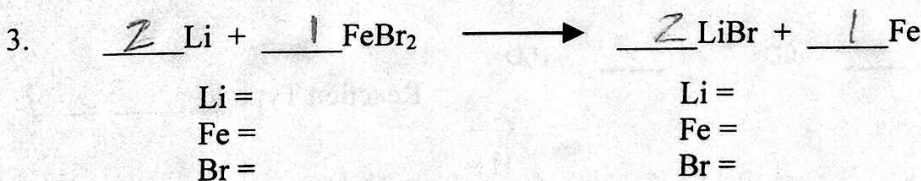
3 USE NOTES TO FILL IN



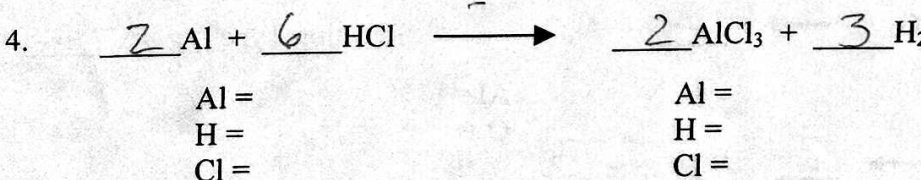
Reaction Type SYNTHESIS



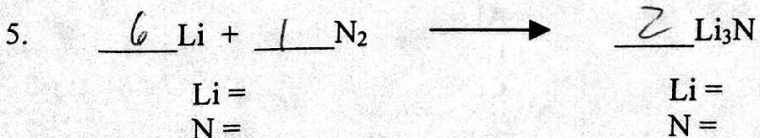
Reaction Type SYNTHESIS



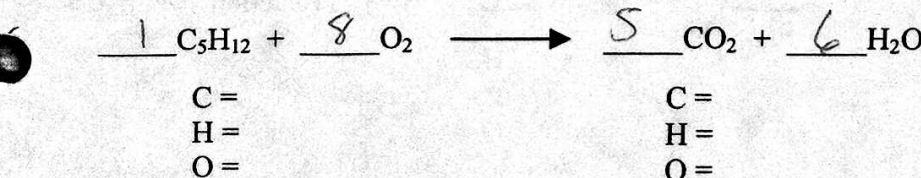
Reaction Type SING REP



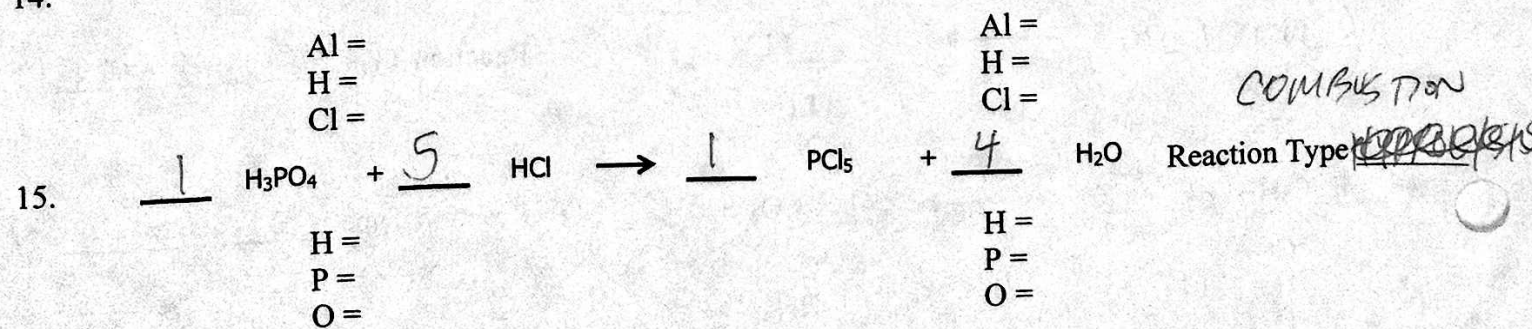
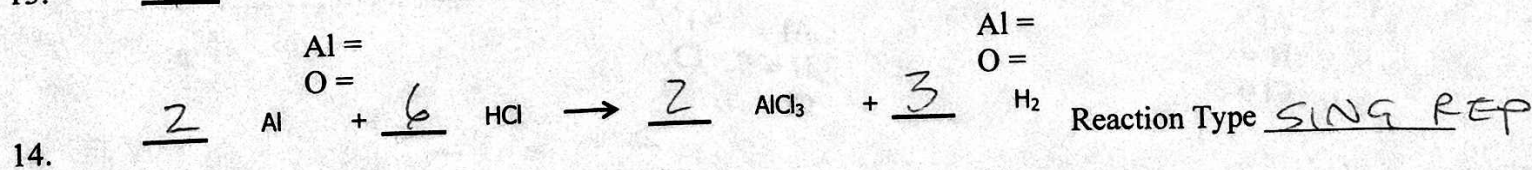
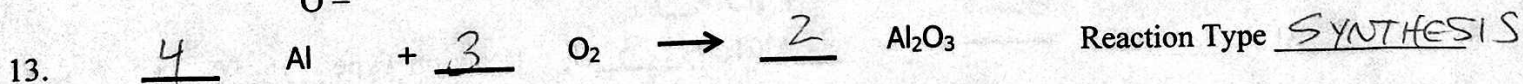
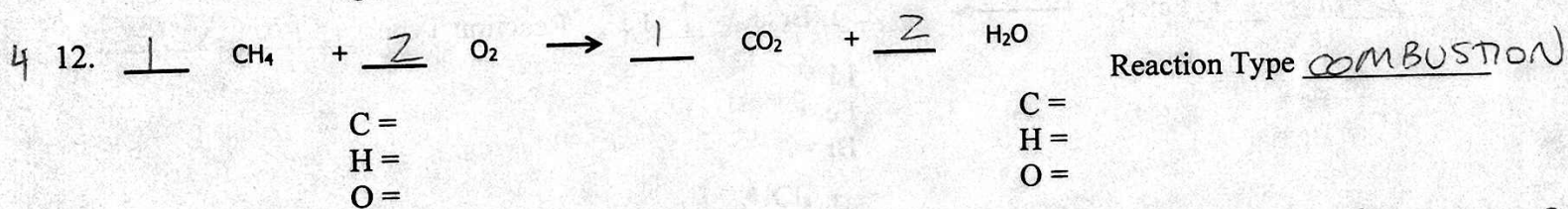
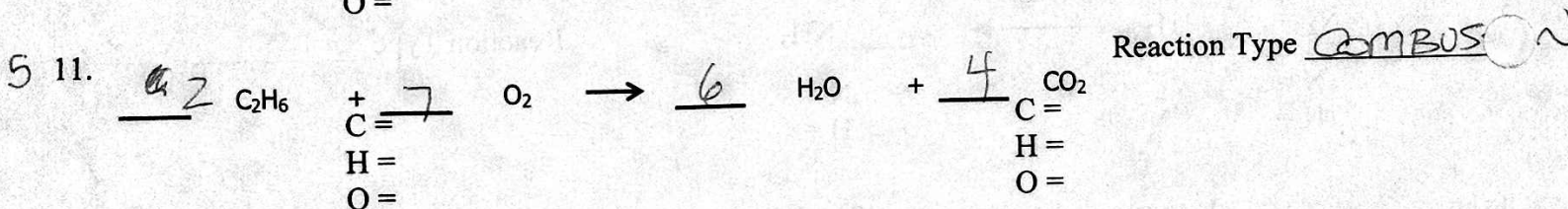
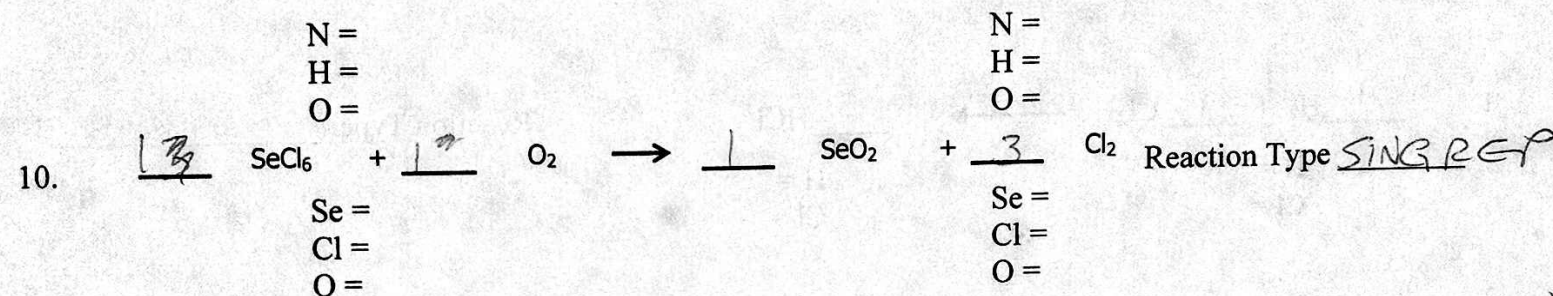
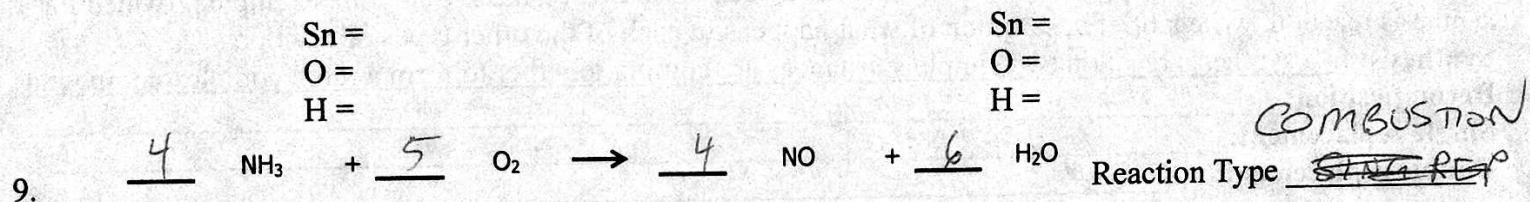
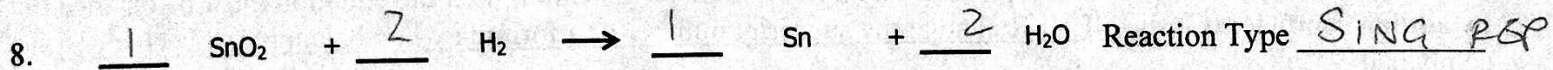
Reaction Type SING REP



Reaction Type SYNTHESIS



Reaction Type COMBUSTION



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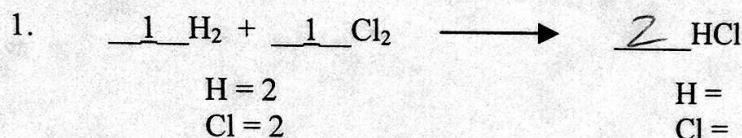
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Decomposition: _____

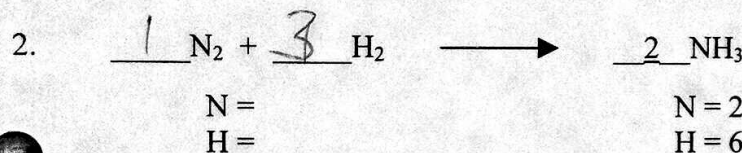
Single replacement: _____

Double replacement: _____

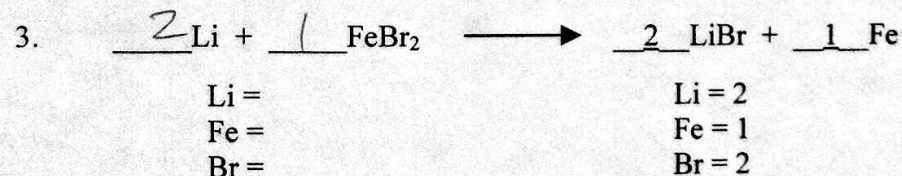
USE NOTES TO FILL IN



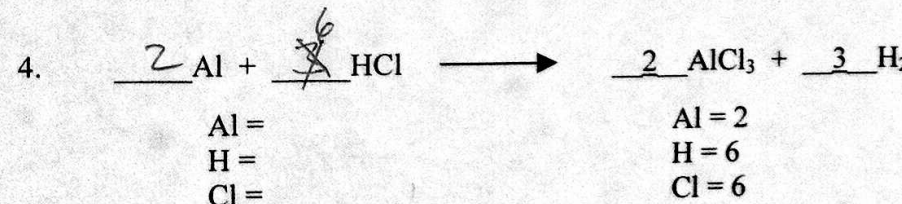
Reaction Type SYNTHESIS



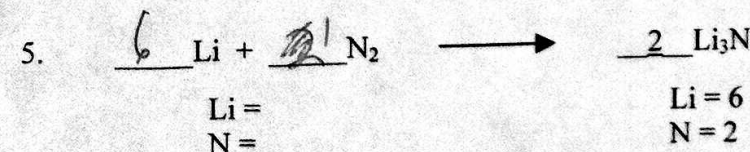
Reaction Type SYNTHESIS



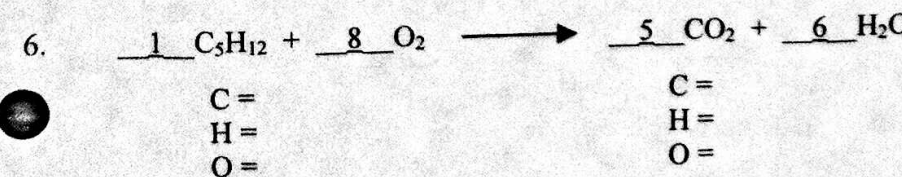
Reaction Type SING REP



Reaction Type SING REP



Reaction Type SYNTHESIS

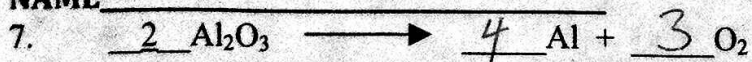


Reaction Type COMBUSTION

PHYSICAL SCIENCE

NAME _____

DSHS
MRS. ELLIS



Al = 4
O = 6

Al =
O =

Reaction Type 34. DECOMPOS

IRON



Na =

Na =

Cl =

Cl =



Al =

Al =

O =

O =

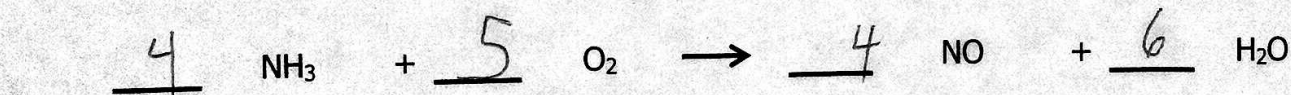


N =

N =

H =

H =



N =

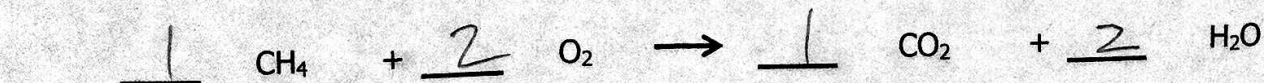
N =

H =

H =

O =

O =



C =

C =

H =

H =

O =

O =